

Continue

Polyatomic Ions Worksheet Answer Key

Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

sulfate	<u>SO₄²⁻</u>	<u>CO₃²⁻</u>	carbonate
nitrite	<u>NO₂⁻</u>	<u>MnO₃¹⁻</u>	manganate
perphosphate	<u>PO₄³⁻</u>	<u>SO₅²⁻</u>	persulfate
hypoiodite	<u>IO⁻</u>	<u>BrO₂¹⁻</u>	bromite
chlorite	<u>ClO₂</u>	<u>CO₄²⁻</u>	percarbonate
phosphite	<u>PO₃³⁻</u>	<u>PO₅³⁻</u>	perphosphate
percarbonate	<u>CO₄²⁻</u>	<u>ClO¹⁻</u>	hypochlorite
bromate	<u>BrO₃</u>	<u>IO₃¹⁻</u>	iodite
hyposulfite	<u>SO₂³⁻</u>	<u>PO₄³⁻</u>	phosphate
permanganate	<u>MnO₄</u>	<u>NO₂¹⁻</u>	nitrite
carbonite	<u>CO₂²⁻</u>	<u>SO₄²⁻</u>	sulfate

2. Name or write the formula for the following Type I polyatomic ionic compounds

beryllium hydroxide	<u>Be(OH)₂</u>	<u>Ba(IO₃)₂</u>	barium iodate
sodium nitrite	<u>NaNO₂</u>	<u>Ga(CNO)₃</u>	gallium cyanate
ammonium chloride	<u>NH₄Cl</u>	<u>Ag₂SO₃</u>	silver sulfite
calcium bisulfate	<u>Ca(HSO₄)₂</u>	<u>MgCO</u>	magnesium hypocarbonite
rubidium perchlorate	<u>RbClO₄</u>	<u>NH₃NO₂</u>	ammonium nitrite
strontium sulfite	<u>SrSO₃</u>	<u>Al(C₂H₃O₂)₃</u>	aluminum acetate
aluminum acetate	<u>Al(C₂H₃O₂)₃</u>	<u>SrSO₅</u>	strontium persulfate
ammonium nitrate	<u>NH₄NO₃</u>	<u>RbClO₂</u>	rubidium chlorite
magnesium hypocarbonite	<u>MgCO</u>	<u>Ca(HSO₄)₂</u>	calcium bisulfate
silver hyposulfite	<u>Ag₂SO₂</u>	<u>NH₄Cl</u>	ammonium chloride

Ionic Formula Worksheet Answer Key

IONS	nitrate	sulfate	carbonate	phosphate	hydroxide	chromate
sodium	NaNO ₃	Na ₂ SO ₄	Na ₂ CO ₃	Na ₃ PO ₄	NaOH	Na ₂ CrO ₄
silver	AgNO ₃	Ag ₂ SO ₄	Ag ₂ CO ₃	Ag ₃ PO ₄	AgOH	Ag ₂ CrO ₄
ammonium	NH ₄ NO ₃	(NH ₄) ₂ SO ₄	(NH ₄) ₂ CO ₃	(NH ₄) ₃ PO ₄	NH ₄ OH	(NH ₄) ₂ CrO ₄
mercury(I)	Hg ₂ (NO ₃) ₂	Hg ₂ SO ₄	Hg ₂ CO ₃	(Hg ₂) ₃ (PO ₄) ₂	Hg ₂ (OH) ₂	Hg ₂ CrO ₄
zinc	Zn(NO ₃) ₂	ZnSO ₄	ZnCO ₃	Zn ₃ (PO ₄) ₂	Zn(OH) ₂	ZnCrO ₄
calcium	Ca(NO ₃) ₂	CaSO ₄	CaCO ₃	Ca ₃ (PO ₄) ₂	Ca(OH) ₂	CaCrO ₄
magnesium	Mg(NO ₃) ₂	MgSO ₄	MgCO ₃	Mg ₃ (PO ₄) ₂	Mg(OH) ₂	MgCrO ₄
copper(I)	CuNO ₃	Cu ₂ SO ₄	Cu ₂ CO ₃	Cu ₃ PO ₄	CuOH	Cu ₂ CrO ₄
lead(II)	Pb(NO ₃) ₂	PbSO ₄	PbCO ₃	Pb ₃ (PO ₄) ₂	Pb(OH) ₂	PbCrO ₄
aluminum	Al(NO ₃) ₃	Al ₂ (SO ₄) ₃	Al ₂ (CO ₃) ₃	AlPO ₄	Al(OH) ₃	Al ₂ (CrO ₄) ₃
manganese(III)	Mn(NO ₃) ₃	Mn ₂ (SO ₄) ₃	Mn ₂ (CO ₃) ₃	MnPO ₄	Mn(OH) ₃	Mn ₂ (CrO ₄) ₃
cobalt(III)	Co(NO ₃) ₃	Co ₂ (SO ₄) ₃	Co ₂ (CO ₃) ₃	CoPO ₄	Co(OH) ₃	Co ₂ (CrO ₄) ₃
copper (II)	Cu(NO ₃) ₂	CuSO ₄	CuCO ₃	Cu ₃ (PO ₄) ₂	Cu(OH) ₂	CuCrO ₄
iron (III)	Fe(NO ₃) ₃	Fe ₂ (SO ₄) ₃	Fe ₂ (CO ₃) ₃	FePO ₄	Fe(OH) ₃	Fe ₂ (CrO ₄) ₃
lead (IV)	Pb(NO ₃) ₄	Pb(SO ₄) ₂	Pb(CO ₃) ₂	Pb ₃ (PO ₄) ₄	Pb(OH) ₄	Pb(CrO ₄) ₂
potassium	KNO ₃	K ₂ SO ₄	K ₂ CO ₃	K ₃ PO ₄	KOH	K ₂ CrO ₄
barium	Ba(NO ₃) ₂	BaSO ₄	BaCO ₃	Ba ₃ (PO ₄) ₂	Ba(OH) ₂	BaCrO ₄

KEY

1. Name or write the formula for the following polyatomic ions

2. Name or write the formula for the following Type I polyatomic ionic compounds

3. Name or write the formula for the following polyatomic ionic compounds

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